

Antioxidant Activity of the Aqueous Extract of *Iris taochia* and Identification of its Chemical Constituents

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Askin *et al.*: Antioxidant activities of aqueous extract from *Iris taochia*

In this study, phytochemical components and antioxidant capacity of extracts obtained from *Iris taochia* plant, which is endemic to Caucasus, were investigated. Fifteen new compounds were detected and identified using a GC-MS system, which were isolated from *Iris taochia* for the first time. From the calibration curve, the amount of α -amyrin in dry powder of *Iris taochia* was calculated. The mean amount of α -amyrin found in whole plant powder of *Iris taochia* was 4.598 $\mu\text{g}/10$ mg with a percent recovery of 0.045 % for α -amyrin. Antioxidant properties of *Iris taochia*, *in vitro* using Fe^{3+} - Fe^{2+} , Cu^{2+} - Cu^{+} and Fe^{3+} -2,4,6-tripyridyl-s-triazine reducing, 1,1-diphenyl-2-picrylhydrazyl scavenging, 2,2'-azino-bis(3-ethylbenzothiazoline-6-sulphonic acid) scavenging, N,N-dimethyl-p-phenylenediamine scavenging and Fe^{2+} chelating activities were measured. The antioxidative effects of *Iris taochia* have not been reported earlier. These results shown that crude ethanol and water extracts of *Iris taochia* had effective reducing effect, radical scavenging activity and metal chelating effects when compared to standard antioxidants. The results indicated that the ethanol and water extracts of *Iris taochia* could be used as a potential source of natural antioxidants by the pharmaceutical industry or as a conceivable food supplement.

Key words: Antioxidant activity, reducing power, α -amyrin, GC-MS, *Iris taochia*

Reactive oxygen species (ROS) including free radicals such as hydroxyl radicals (OH^{\bullet}), singlet oxygen ($^1\text{O}_2$), superoxide anion radicals ($\text{O}_2^{\bullet-}$), and non-free radical species such as hydrogen peroxide (H_2O_2) are diverse kinds of activated oxygen and mostly produced by exogenous factors or oxidation product of biological reactions^[1-3]. ROS have created considerable interest among scientists in the recent years^[4-6]. Their wide range of effects that are of medicinal and biological interest have led to several experimental studies^[7-9]. Additionally, it is well-known that ROS induce oxidative damage to biomolecules such as lipids, proteins, carbohydrates and deoxyribonucleic acids^[10-12], which might induce cancer, ageing, and other multitude of diseases^[13-16]. In addition, ROS has also been implicated in more than 100 diseases such as heart disease, acquired immunodeficiency syndrome, malaria, stroke, diabetes, arteriosclerosis, and cancer^[17-20].

The plant kingdom, including medicinal and dietary plants, offers many natural phytochemicals that include phenolic diterpenes, triterpenes, flavonoids,

phenolic acids and sterols^[21-23]. Antiinflammatory, antioxidant and anticancer activities reported to be possessed by these compounds might have resulted from prevention of oxidative damage^[24,25]. There is rising interest in endemic species because these can be used for preparing phytopharmaceuticals with considerable antioxidant potential or for the production of raw materials and health confidants^[26-28]. Antioxidant compounds are molecules that can inhibit or delay the oxidation of lipids or other biomolecules by inhibiting the initiation and propagation of oxidative chain reactions^[29]. Some plants synthesize large amounts of vitamin C, vitamin E, and carotenoids, which are well-known antioxidants^[30-33]. Phenolic compounds are also widely distributed in plants and have been found to possess antioxidative potential as well^[34]. Natural

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antioxidant compounds can retard the progression of chronic diseases and protect human cells from free radicals as well as prevent oxidative rancidity of lipid containing foods^[35-37].

Iris taochia is endemic to the Caucasus. This plant is quite rich in triterpenes and sterols^[38]. Terpenes play significant role as growth regulators (phytohormones) of plants and signal compounds, secondary pigments pending photosynthesis and also with antimicrobial activities^[39]. Terpene compounds include cucurbitacins, cardenolides, bufadienolides, phytosterols, sesquiterpenoids, iridoids, diterpenoids, sesquiterpene lactones, triterpenoid saponins, steroid saponins, carotenoids, nortriterpenoids, other triterpenoids, and monoterpenoids^[40]. Triterpene saponins are the most valued in terms of pharmacology^[41]. These compounds have been reported to possess antiinflammatory, hypoglycemic and anticancer activities^[42]. The amyryns are three closely related natural compounds of the triterpene class^[43]. They are designated α -amyryn (ursane skeleton), β -amyryn (oleanane skeleton) and δ -amyryn^[42,43]. In this investigation chemical constituents and antioxidant potential of dried whole plant of *Iris taochia* was studied.

It appeared from the literature reports, the phenolic constituents and the oxidative potential of *I. taochia* has not been reported. The antioxidant potential of *I. taochia* was investigated using seven different antioxidant procedures including ferric reducing antioxidant power (FRAP), cupric ion reducing antioxidant capacity (CUPRAC), Fe^{3+} - Fe^{2+} reducing, 2,2'-azino-bis-3-ethylbenzothiazoline-6-sulphonic acid (ABTS⁺) scavenging, N,N-dimethyl-p-phenylenediamine (DMPD⁺) scavenging, Fe^{2+} chelating, and 1,1-diphenyl-2-picrylhydrazyl (DPPH[•]) assays. On the other hand, the antioxidant effect of the extract was decisively related to the total flavonoid and phenolic contents of the extract that considered in this paper.

MATERIAL AND METHODS

Collection and extraction of *Iris taochia*:

Aerial parts and rhizomes of *I. taochia* were collected from Tortum in the province of Erzurum in April to June 2016 and were authenticated as per the characteristics described by Davis *et al.* and Richardson^[44,45]. Specimens were air-dried at room temperature for 7 d and stored for analysis at a later date. The drying process of the plant and all the remaining studies were

carried out in the Genetics Laboratory, Science Faculty, Ataturk University. The air-dried and powdered aerial parts and rhizomes of the plant were extracted thrice with ethanol in a mantle heater at 40°. Ethanol extraction was also carried out according to the method of Kotan *et al.*^[46]. In this method, 100 g dried plant powder was shaken with 2 times the volume of ethanol on a shaker for 3 days at room temperature. At the end of d 3, the extract was filtered, 1 μl of the filtrate was employed for GC-MS analysis. This sample was also used in antioxidant capacity assays.

GC-MS system:

Chromatographic analysis was carried out on an Agilent 7820A gas chromatography system equipped with 5977 series Mass Selective Detector, 7673 series Autosampler and Chemstation (Agilent Technologies, Palo Alto, CA). HP-5 MS column with 0.25 μm film thickness (30 m \times 0.25 mm I.D., USA) was used for separation. The temperatures of the inlet, transfer line and detector were 250, 250 and 300°, respectively.

GC-MS conditions:

Diverse temperature programs were evaluated for GC-MS procedure. Towards the end of this evaluation, the temperature program chosen was as follows, initial temperature was 50° held for 1 min, increased to 100° at a rate of 20°/min, held for 1 min, increased to 180° at a rate of 10°/min, held for 1 min, increased to 220° at a rate of 5°/min, held for 5 min, and finally to 300° at a rate of 10°/min and held for 5.5 min. The injector volume was 1 ml in splitless mode and the carrier gas was helium at a flow rate of 1 ml/min.

Antioxidant capacity assays

The *in vitro* antioxidant potential of *I. taochia* was evaluated using various procedures that included, cupric ion (Cu^{2+}) reducing power (CUPRAC method), FRAP reduction of Fe^{3+} -2,4,6-Tris(2-pyridyl)-s-triazine (Fe^{3+} -TPTZ) complex, DPPH[•] scavenging, ferrous ions (Fe^{2+}) chelating, DMPD radical scavenging and ABTS⁺ scavenging activities. Additionally, butylated hydroxytoluene (BHT), butylated hydroxyanisole (BHA), α -tocopherol and trolox (6-hydroxy-2,5,7,8-tetramethylchroman-2-carboxylic acid) were used as standard antioxidants.

Fe^{3+} reducing power assay:

Reducing power was determined by the straight reduction of $\text{Fe}^{3+}(\text{CN}^-)_6$ to $\text{Fe}^{2+}(\text{CN}^-)_6$, which was obtained by evaluating absorbance resulted from the organization of the Perl's Prussian blue complex

following the addition of excess ferric ions (Fe^{3+}). This method is dependent on the reduction of (Fe^{3+}) ferricyanide in stoichiometric overplus relative to the antioxidants. Briefly, different concentrations of this plant extract (10-30 $\mu\text{g/ml}$) in 0.75 ml of deionized H_2O were added with 1.25 ml of phosphate buffer (0.2 M, pH 6.6) and 1.25 ml of potassium ferricyanide ($\text{K}_3\text{Fe}(\text{CN})_6$, 1 %). Then, the solution was incubated at 50° for 20 min. After the incubation, 1.25 ml of trichloroacetic acid (TCA, 10 %) was added. Lastly, 0.5 ml of FeCl_3 (0.1 %) was transferred to this mixture and the absorbance was measured at 700 nm in a spectrophotometer^[36].

CUPRAC assay:

The reducing capacity of the extract was determined by the CUPRAC procedure^[4]. In this method, 0.25 ml of $\text{CH}_3\text{COONH}_4$ buffer solution (1.0 M), 0.25 ml of neocuproine solution in ethanol (7.5×10^{-3} M), 0.25 ml CuCl_2 solution (0.01 M) were mixed and the sample extract at various concentrations (10-50 g/ml) was added to this mixture. The final volume was adjusted to 2 ml with distilled water and 30 min later absorbances of the samples were measured at 450 nm. Increase in absorbance was defined as increase in reducing capacity.

FRAP assay:

FRAP assay was based on reduction of Fe^{3+} -TPTZ complex under acidic conditions. Increased absorbance of blue-coloured ferrous form (Fe^{2+} -TPTZ complex) was recorded at 593 nm. TPTZ solution (2.25 ml, 10 mM TPTZ in 40 mM HCl) was freshly prepared, then transferred to acetate buffer (25 ml, 0.3 M, pH 3.6), and FeCl_3 solution (2.25 ml, 20 mM) in water. Then, different concentrations of the plant extract (10-50 $\mu\text{g/ml}$) were dissolved in 5 ml of appropriate buffer, stirred and incubated at 37° for 30 min. Finally the absorbance of the mixture was measured at 593 nm^[36].

DPPH[•] scavenging activity:

Electron or hydrogen donating capabilities of the samples was determined using DPPH method^[47]. Accordingly, purple-colored 1 mM DPPH solution in ethanol was added to samples at varying concentrations (10-50 $\mu\text{g/ml}$). Finally, the mixture was incubated in a dark room for 30 min and the radical scavenging ability was determined spectrophotometrically at 517 nm against a blank^[46]. Reduced absorbance of the sample indicated the DPPH free radical scavenging capability.

ABTS^{•+} scavenging activity:

ABTS radical cation was generated by the interaction of ABTS (7 mM) and $\text{K}_2\text{S}_2\text{O}_8$ (2.45 mM) as described previously^[47]. One milliliter of ABTS^{•+} solution was added to 3 ml of novel symmetric sulfamides and control solutions. ABTS^{•+} was produced by the reaction of 2 mM ABTS in H_2O with 2.45 mM potassium persulfate ($\text{K}_2\text{S}_2\text{O}_8$) stored in the dark at room temperature for 4 h. Before usage, the ABTS^{•+} solution was diluted to get an absorbance of 0.750 ± 0.025 at 734 nm with sodium phosphate buffer (0.1 M, pH 7.4). Then, to 1 ml of ABTS^{•+} solution, 3 ml of *I. taochia* solution in ethanol at different concentrations (10-50 $\mu\text{g/ml}$) was added. After 30 min, the percent inhibition of ABTS at 734 nm was calculated for each concentration relative to a blank absorbance^[47]. The extent of decoloration was calculated as percent reduction of absorbance.

Metal chelating assay:

Ferrous ions (Fe^{2+}) chelating activities of *I. taochia* were investigated conforming to the method described previously^[48]. Different concentrations (10-50 $\mu\text{g/ml}$) of this plant extract in 0.25 ml FeSO_4 solution (2 mM), 0.25 ml ethanol, 1 ml ferrozine solution (0.2 % in 0.2 M HCl), 1 ml Tris-HCl buffer solution (pH 7.4), and 2.5 ml ethanol were placed into a test tube. The absorbance was measured at 562 nm.

DMPD^{•+} scavenging activity:

DMPD^{•+} scavenging capability of *I. taochia* extract was obtained conforming to Fogliano *et al.*^[49], as previously reported^[26]. DMPD (100 mM) was prepared by dissolving 209 mg of DMPD in 10 ml of deionized water and 1 ml of this solution was added to 100 ml of 0.1 M acetate buffer (pH 5.3) and the colored radical cation (DMPD^{•+}) was obtained by adding 0.2 ml of a solution to 0.05 M ferric chloride (FeCl_3). The absorbance of this solution, which was freshly prepared daily, is constant up to 12 h at room temperature. Different concentrations of standard antioxidants or resveratrol (10-50 $\mu\text{g/ml}$) were added in test tubes and the total volume was adjusted with distilled water to 0.5 ml^[26]. Ten minutes later, the scavenging capability of DMPD^{•+} radical of the sample was spectrophotometrically measured at 505 nm.

Determination of total phenolic content by the Folin-Ciocalteu assay:

The total phenolic contents of *I. taochia* was obtained using the Folin-Ciocalteu method^[50], as previously described^[51]. Initial stock solutions of *I. taochia* at

2 mg/ml were prepared. Absorbance (λ_{760}) = $0.0016 \times \text{total phenols (gallic acid equivalent, GAE; } \mu\text{g)}$. The content of total phenolics in *I. taochia* was determined by employing this graph (R^2 : 0.9872), which was prepared using gallic acid and expressed as μg of GAE. Briefly, 1 ml of extract containing 1 g extract was mixed with 45 ml distilled water. One millilitre of Folin-Ciocalteu reagent was added and the content of the flask mixed thoroughly. After 3 min, 3 ml of 2 % Na_2CO_3 was added, the mixture was allowed to stand for 2 h with intermittent shaking and the absorbance was measured at 760 nm.

Determination of total flavonoid contents:

The total flavonoid content of *I. taochia* was approximately obtained by a colorimetric method, which was reported previously^[52]. One milligram of the sample was added into a test tube. Then 0.1 ml CH_3COOK (1.0 M) and 0.1 ml of 10 % $\text{Al}(\text{NO}_3)_3$ in 4.3 ml ethanol were added and the samples were vortexed. The vortexed samples were kept at room temperature for 40 min. The absorbance of the samples was recorded at 415 nm. Results were reported as μg quercetin equivalents (QE) per mg extract. A quercetin calibration curve was prepared, and the amount of flavonoid was calculated using the linear regression

equation that was determined from the calibration curve. Finally, the results were expressed as QEs/mg extract. Absorbance (λ_{415}) = $0.011 \times \text{total phenols (QE } \mu\text{g)}$. The total flavonoid content in *I. taochia* was calculated using the graph, which was prepared using quercetin and determined as QE μg .

RESULTS AND DISCUSSION

In this investigation, fifteen compounds were obtained and identified as L- α -terpineol, tetradecanoic acid, neophytadiene, 2-pentadecanone, 6,10,14-trimethyl phthalic acid, n-hexadecanoic acid, linoleic acid, linolenic acid, oleic acid, octadecanoic acid (stearic acid), hexacosanoic acid, docosanoic acid 1,2,3-propanetriylester, β -sitosterol, α -tocopherol and α -amyrin (fig. 1 and Table 1).

To validate the present method, parameters such as linearity, precision, accuracy, limit of detection (LOD), limit of quantification (LOQ) and recovery were investigated according to International Conference on Harmonization validation guidelines (ICH, 1996). The linear range of α -amyrin in this developed method was 1-100 $\mu\text{g/ml}$. The intra- and inter-day precisions, expressed as the relative standard deviation (RSD), were less than 4.97 and 4.99 %, determined from quality

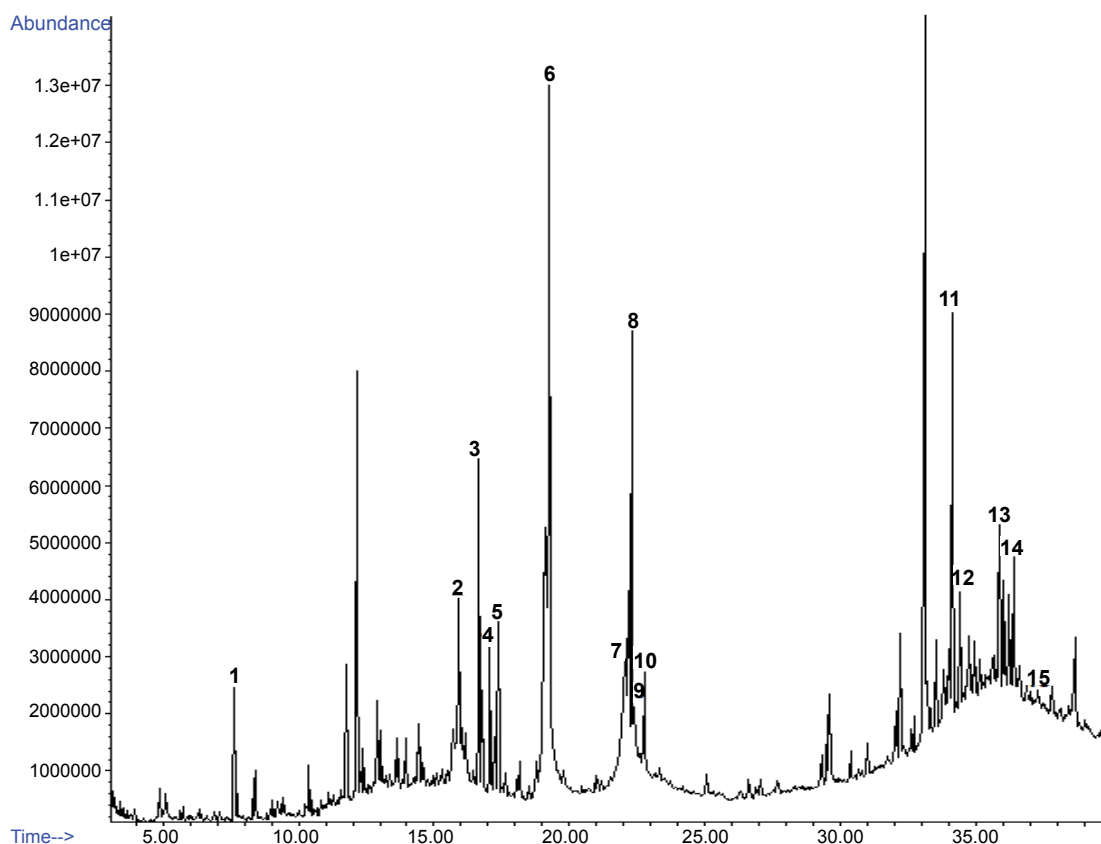
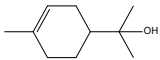
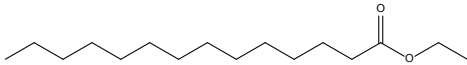
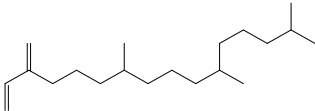
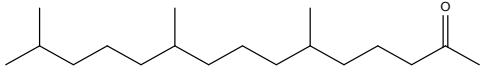
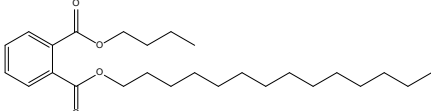
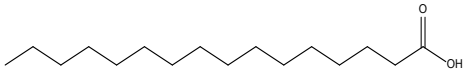
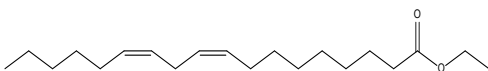
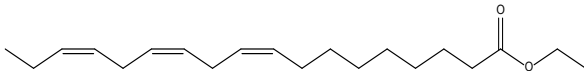
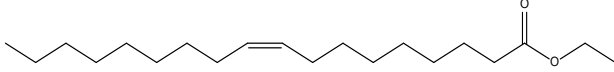
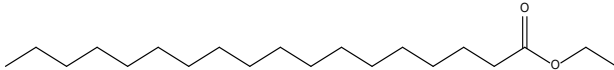
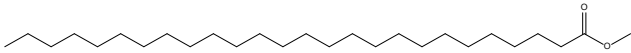
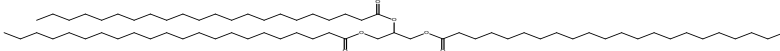
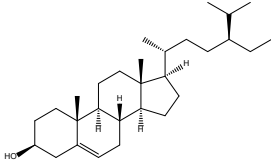
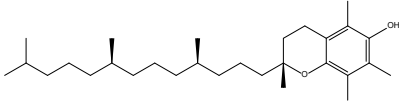
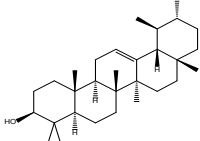


Fig. 1: A typical chromatogram of *I. taochia*

TABLE 1: STRUCTURAL FORMULAE OF COMPOUNDS DETECTED IN WHOLE PLANT OF *I. TAOCHIA*

Compound Name	Chemical structure	Retention Time (min)
L- α -terpineol		7.72
Tetradecanoic acid ethyl ester		15.92
Neophytadiene		16.66
2-Pentadecanone, 6,10,14-trimethyl		16.78
Phthalic acid, butyl tetradecyl ester		17.24
n-Hexadecanoic acid		19.1
Linoleic acid ethyl ester		22.15
Linolenic acid ethyl ester		22.29
Oleic acid ethyl ester		22.37
Octadecanoic acid ethyl ester		22.75
Hexacosanoic acid methyl ester		34.4
Docosanoic acid 1,2,3-propanetriyl ester		34.9
β -sitosterol		36
α -tocopherol		36.4
α -amyrin		37

control samples for α -amyrin, and accuracy was within 2.00 and 4.60 % in terms of relative error, respectively. The percent recovery obtained for α -amyrin was 99.7 %. LOD and quantification for α -amyrin were 5 and 15 ng/ml, respectively.

Also, the developed method can be used for routine quality control analysis of α -amyrin in dried whole plant of *I. taochia*. The sample working solution (1 μ l) was injected and the height of α -amyrin peak was measured. From the calibration curve, the amount of α -amyrin in dry powder of *I. taochia* was calculated. The mean amount of α -amyrin found in the whole plant powder of *I. taochia* was 4.598 μ g/10 mg with percent recoveries 0.045 % for α -amyrin.

Reducing ability of *I. taochia* and standard compounds was given in (Table 2 and fig. 2). These results demonstrated that crude ethanol extract (EEIT) exhibited reducing effect when compared to standard antioxidants. Also, radical scavenging activity and metal chelating effects of *I. taochia* and standard compounds were summarized in (Table 3 and fig. 2). The IC₅₀ values of the EEIT and lyophilized water extract of *I. taochia* (WEIT) in the DPPH⁺, ABTS⁺, DMPD⁺ and metal chelating assays were 24.75, 18.23, 40.76, 57.75, and 30.13, 23.10, 57.75, 86.62 μ g/ml, respectively. In addition to these, 47.50 μ g GAE phenolic compounds and 10.72 μ g QE flavonoids were detected in 1 mg dried EEIT (Table 4).

The IC₅₀ values for DPPH scavenging by standard antioxidants and both extracts were compared. The IC₅₀ of WEIT was 30.13 μ g/ml, which was higher than that of EEIT, 24.75 μ g/ml, but both these values were much higher than those of Trolox (12.51 μ g/ml), BHT (10.65 μ g/ml), BHA (10.34 μ g/ml), and α -tocopherol (9.86 μ g/ml). The lower the IC₅₀ value the higher the DPPH radical scavenging effect. These results indicated that both EEIT and WEIT possessed DPPH radical scavenging activities. In this investigation, DMPD scavenging, ABTS radical cation scavenging

and Fe chelating effects were also evaluated to detect radical scavenging activity. In these procedures, EEIT and WEIT exhibited similar reduction in free radicals when compared to those obtained in the DPPH reaction. The IC₅₀ values for ABTS radical scavenging of EEIT, WEIT, and standard antioxidants were as follows, WEIT with an IC₅₀ value of 23.10 μ g/ml and EEIT with an IC₅₀ value of 18.23 μ g/ml were slightly higher than those of α -tocopherol, 11.55 μ g/ml, Trolox, 2.88 μ g/ml, BHA, 1.86 μ g/ml, and BHT, 1.75 μ g/ml. In the DPPH, ABTS, and metal chelating tests the antioxidant activities by these two extracts although were less than the standard antioxidants, both EEIT and WEIT could be considered to have effective radical scavenging activity. Additionally, the DMPD scavenging method was also used to estimate detect radical scavenging activity. In this procedure, the IC₅₀ values obtained were in the following order; α -tocopherol (86.62 μ g/ml)>WEIT (57.75 μ g/ml)>BHT (46.20 μ g/ml)>EEIT (40.76 μ g/ml)>BHA (34.65 μ g/ml)>trolox (28.87 μ g/ml). These values indicated that the plant extracts exhibited direct radical scavenging activity within the range of activities exhibited by the standard antioxidants.

Antioxidant molecules induced reduction of the Fe³⁺/ferricyanide complex to its ferrous (Fe²⁺/ferricyanide) form due to their reducing potential. As shown in Table 2, the ferric reducing power increased with increasing concentration of both WETT and EETT, which however were lower than the standard antioxidants. The reducing power of the standard antioxidants and the extracts at 50 μ g/ml were as follows, BHT (2.607, R²: 0.9818) = BHA (2.304, R²: 0.9733)>trolox (2.277, R²: 0.9722)> α -tocopherol (1.744, R²: 0.9602)>EEIT (1.308, R²: 0.9889)>WEIT (1.001, R²: 0.9582).

In this investigation the CUPRAC procedure was also used to get an estimate of the reducing power of the extracts. This procedure is based on the reduction

TABLE 2: REDUCING POWER OF EEIT, WEIT AND STANDARD COMPOUNDS

Antioxidants	Fe ³⁺ -Fe ²⁺ reducing		Cu ²⁺ -Cu ⁺ reducing		Fe ³⁺ -TPTZ reducing	
	λ_{700}	r ²	λ_{450}	r ²	λ_{593}	r ²
BHA	2.304±0.013	0.9733	2.803±0.020	0.9424	3.000±0.017	0.9711
BHT	2.607±0.012	0.9818	2.506±0.024	0.9616	3.000±0.012	0.9636
α -Tocopherol	1.744±0.017	0.9602	1.371±0.019	0.9381	2.627±0.001	0.9844
Trolox	2.277±0.013	0.9722	1.202±0.005	0.9783	2.932±0.015	0.9744
EEIT	1.308±0.012	0.9889	1.638±0.011	0.9905	1.490±0.005	0.9550
WEIT	1.001±0.017	0.9582	1.303±0.009	0.9690	1.288±0.006	0.9937

Reducing activity of 50 μ g/ml of EEIT (ethanol extract of *I. taochia*), WEIT (lyophilized water extract of *I. taochia*) and standards by FRAP assay, ferric ions (Fe³⁺) reducing, and cupric ions (Cu²⁺) reducing assays

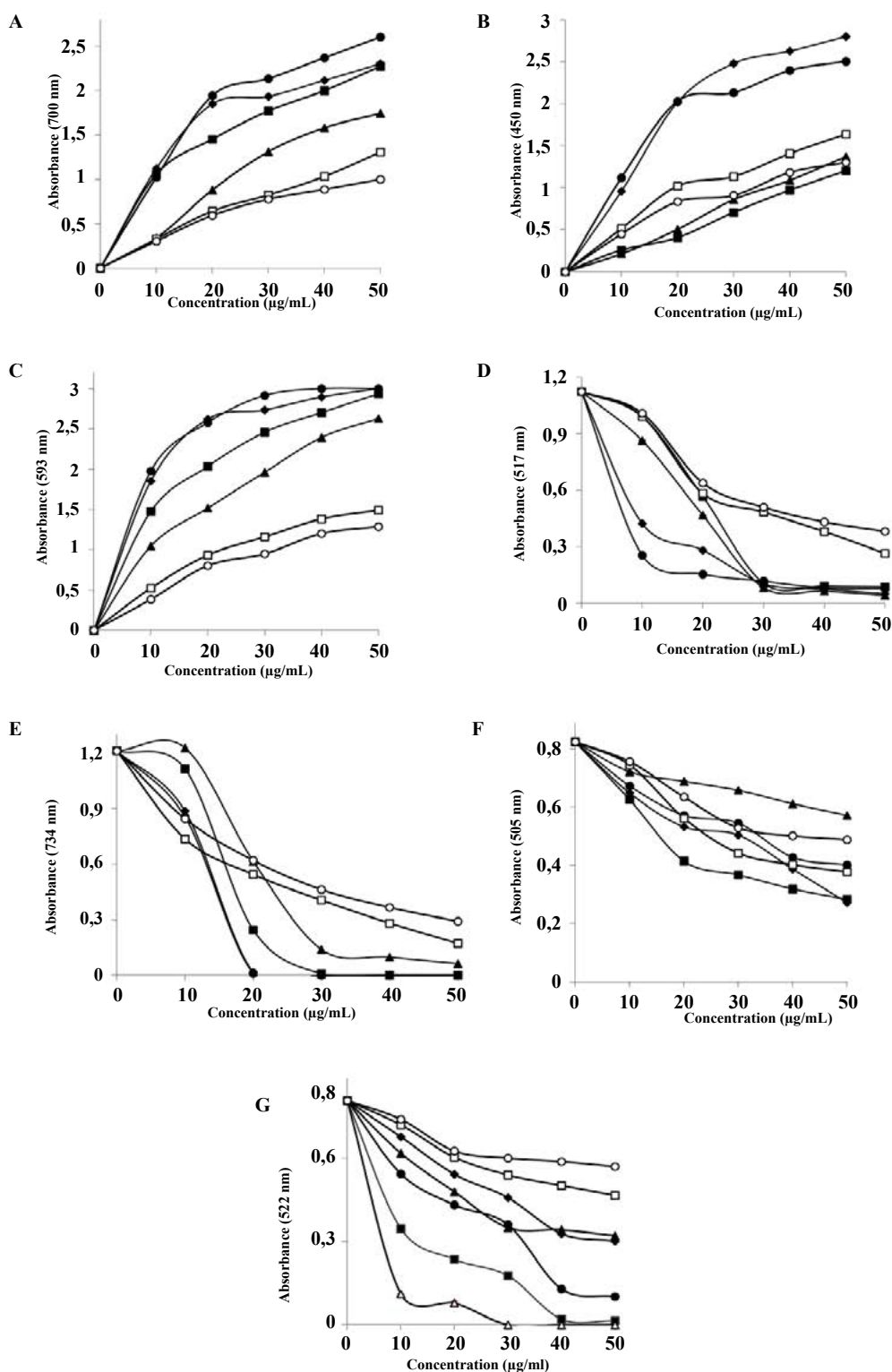


Fig. 2: Reducing and radical scavenging activity of EEIT, WEIT and standard compounds

WEIT: lyophilized water extract of *I. taochia*; EEIT: ethanol extract of *I. taochia*. A. Fe^{3+} - Fe^{2+} reducing activity; B. Cu^{2+} - Cu^{+} reducing activity; C. Fe^{3+} -TPTZ reducing assay, D. DPPH assay, E. ABTS^{+} scavenging assay, F. DMPD^{+} scavenging assay and G. Fe^{2+} chelating activity. \blacklozenge — BHA; \bullet — BHT; \blacktriangle — α -tocopherol; \blacksquare — trolox; \triangle — EDTA; \square — EEIT; \circ — WEIT

of Cu^{2+} to Cu^{+} by antioxidants in the presence of neocuproine^[53]. In this method, a higher absorbance indicated a higher cupric ion (Cu^{2+}) reducing power. The cupric ion reducing powers of 50 $\mu\text{g/ml}$ of

EEIT, WEIT, and the standard molecules were as follows, BHA (2.803, R^2 : 0.9424)>BHT (2.506, R^2 : 0.9616)>EEIT (1.638, R^2 : 0.9905)> α -tocopherol (1.371, R^2 : 0.9381)>WEIT (1.303, R^2 : 0.9690)>trolox

TABLE 3: IC₅₀ OF EEIT, WEIT AND STANDARDS IN DPPH[•], ABTS^{••}, DMPD^{••} RADICAL SCAVENGING AND Fe²⁺ CHELATING ASSAYS

Antioxidants	DPPH [•] Scavenging	ABTS ^{••} Scavenging	DMPD ^{••} Scavenging	Fe ²⁺ Chelating
BHA	10.34	1.86	34.65	34.65
BHT	10.65	1.75	46.20	22.35
α -Tocopherol	9.86	11.55	86.62	31.50
Trolox	12.51	2.88	28.87	14.14
EDTA	-	-	-	1.96
EEIT	24.75	18.23	40.76	57.75
WEIT	30.13	23.10	57.75	86.62

Determination of IC₅₀ in μ g/ml of EEIT (ethanol extract of *I. taochia*), WEIT (lyophilized water extract of *I. taochia*) and standards in DPPH[•], ABTS^{••}, DMPD^{••} radical scavenging and Fe²⁺ chelating assays

TABLE 4: TOTAL PHENOLIC AND FLAVONOID CONTENT OF EEIT AND WEIT

	EEIT	WEIT
Total phenols*	47.50	39.26
Total flavonoids**	10.72	9.12

WEIT: lyophilized water extract of *I. taochia*; EEIT: ethanol extract of *I. taochia*. * Determined as μ g of gallic acid equivalent (GAE) in mg dried extracts. ** Determined as μ g of quercetin equivalent (QE) in mg dried extracts

(1.202, R²: 0.9783; Table 2). In addition, the FRAP procedure was also used to further understand the reducing power of the molecules. The reducing power of 50 μ g/ml of EEIT, WEIT, and the standards were as follows, BHA (3.000, R²: 0.9711)=BHT (3.000, R²: 0.9636)>trolox (2.932, R²: 0.9744)> α -tocopherol (2.627, R²: 0.9844)>EEIT (1.490, R²: 0.9550)>WEIT (1.288, R²: 0.9937; Table 2).

The compounds listed in Table 1 were identified in *I. taochia* for the first time. There is no work about the chemical composition and antioxidative activity of *I. taochia* in literature. Therefore, the current study became very significant. These findings might shed light on the design of new drugs. These results indicated that EEIT and WEIT could be very useful to pharmaceutical industry or as a food supplement and as an easily attainable source of natural antioxidants. Additionally, this study demonstrated that *I. taochia* has powerful antioxidant, reducing power, free radical scavenging and metal chelating activities comparable to those possessed by diverse standards such as α -tocopherol, trolox, BHA, BHT. More recently research in food science and nutrition has focused on plant products that possessed antimicrobial and potential antioxidant activities^[54-56]. Antioxidant properties, particularly radical scavenging are of high value due to the detrimental role of free radicals in biological systems and in foods^[57-61]. DPPH[•] assay provided data on the reactivity of test molecules with constant free radicals^[62]. Due to the odd electron quantity, in visible spectroscopy, DPPH[•] and ABTS^{••}

gave powerful absorption bands at 517 and 734 nm, respectively. If the electron pairs off in the presence of a free radical scavenger, the absorption vanishes, and the resulting decolorization is stoichiometric in respect to the number of electrons taken up^[63-66].

Metal chelating capacity of *I. taochia* was considerable, since it decreased the concentration of the hydrolysing transition metal. It was recorded that chelating factors are efficacious as secondary antioxidants due to the decrease the redox potential, finally stabilizing the oxidized form of the metal ion. Fe³⁺-Fe²⁺ transformation was evaluated to obtain the measurements of the reductive capability of *I. taochia* by using the procedure of Oyaiz^[67]. Its results on the reducing power show the electron donor confidants of *I. taochia*, finally neutralizing free radicals by forming fixed products. The effect of the reducing reaction is to finish the radical chain reactions that can otherwise be very harmful^[68,69]. Free radicals are significantly accessible in living systems; although, high contents of free radicals can oxidise biomolecules, leading to tissue harm, degenerative factors or cell death, inflammation containing cancer, aspects of ageing, skin irritations arteriosclerosis, cardiovascular diseases, and neural disorders^[70,71]. Also, the antioxidant activity of the extract was decisively dependent on the total phenolic and flavonoid contents of the extract that determined in this study.

Previous studies were designed to investigate antioxidant and anticholinesterase potential of *I. germanica* var. *florentina*. Acetylcholinesterase (AChE) and butyrylcholinesterase (BChE) inhibitory potential of plant samples were investigated using the Ellman's assay. Antioxidant activity was performed using DPPH, H₂O₂ and ABTS free radical scavenging assays. Total phenolics and flavonoids contents were expressed in mg GAE/g dry weight and mg RTE/g, respectively^[72]. And also, Basgedik *et al.*^[73] investigated the antimicrobial, antioxidant and

antimutagenic properties of ethanolic extracts of the aerial parts and rhizomes of *I. albicans* Lange. While previous studies have examined the antimicrobial and antioxidant properties of *I. albicans*, to our knowledge, this is the first study to report on the antimutagenic activity of this plant. Both aerial part and rhizome extracts exhibited limited antimicrobial activity against *Bacillus subtilis* ATCC 6633. IC₅₀ values for the radical scavenging activity of the extracts were 8.8 and 11.1 mg/ml, respectively. Total antioxidant activity of the extracts (at 3.15 mg/ml) was 96.6±0.07 and 97.2±0.7 %, respectively^[73].

Conforming to data of this paper, when compared to standard antioxidant molecules such as BHT and BHA, the natural antioxidant trolox (a water-soluble analogue of tocopherol) and α -tocopherol, WEIT and EEIT were shown to be effective antioxidants in various *in vitro* methods, including reducing power, DPPH[•] radical, ABTS^{•+} radical scavenging; DMPD scavenging; and metal chelating activities. Also, this study revealed that WEIT and EEIT appear to be excellent sources of antioxidants for medicines, food, and pharmaceuticals.

Conflicts of interest:

There are no conflicts of interest.

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